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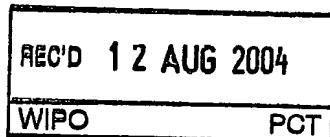
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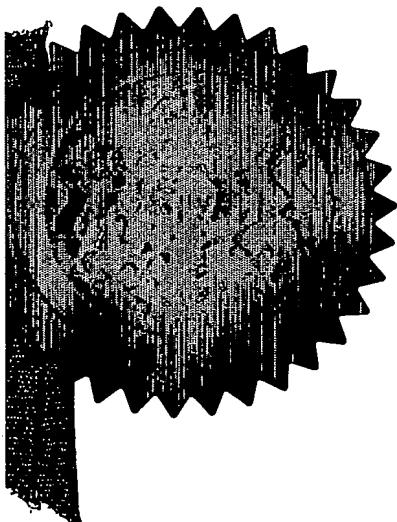


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CHELMSFORD  
ESSEX, CM1 2QU

08457449001

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Reddie & Grose  
16 Theobalds Road  
LONDON  
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DUPLICATE

- 1 -

SENSOR

FIELD OF THE INVENTION

This invention relates to microelectrochemical reaction chambers, and in particular, to a method and apparatus for detecting adverse reaction conditions for microelectrochemical reactions.

BACKGROUND OF THE INVENTION

A particular area of interest in the pharmaceutical industry is the prediction of how drugs are metabolised in the body. For example, one particular parameter of interest is the maximal rate of reaction between a compound and a metabolic enzyme. During a typical metabolism process, an oxidative drug-metabolising enzyme (DME) acts to add a hydroxyl moiety to foreign molecules, thus facilitating their metabolic degradation. The reactions catalysed by enzymes such as DMEs are driven by the transfer of electrons. In particular, an important part of the reaction process is that electrons are supplied to the catalytic site within the enzyme so that the enzyme may react with the compound. Reactions involving the transfer of electrons from one species to another are known as oxidation-reduction (redox) reactions. Usually, the electrons which drive metabolic reactions are supplied *in vivo* by redox partners with the aid of appropriate oxidoreductase enzymes. However, it is known that redox reactions may also be driven artificially by supplying electrons directly using electrodes in an electrochemical reaction chamber. It is in this way that metabolic processes may be studied.

Microelectrochemical reaction chambers are well known in the art. In a typical chamber, a mixture of enzyme, substrate, and mediator is electrochemically reacted by passing an electric current through the mixture, the electric current being provided by a pair of electrodes. The resulting reaction may then be analysed by any suitable means. It is important that the correct conditions are present within the reaction chamber otherwise the properties of any reactions that occur may be affected. In such a case, the results of any analysis of a reaction may be erroneous. One way in which the reaction conditions may be adverse is if the concentrations of reagents in a reaction mixture are altered. Another way in which the reaction conditions may be affected is by the presence of bubbles within the reaction chamber.

We have appreciated the importance of detecting adverse reaction conditions so that erroneous analysis may be rejected. We have further appreciated that a means to detect adverse reaction conditions may be advantageously provided by existing components of reaction apparatus.

dc.

#### SUMMARY OF THE INVENTION

The invention is defined in the independent claims, to which reference may now be directed. Advantageous features are set out in the dependant claims.

5 In a broad aspect, an embodiment of the invention provides a method and apparatus for detecting that conditions in a microelectrochemical reaction chamber are adverse. In use, a voltage is applied by electrodes across a region of reaction mixture located within the reaction 10 chamber. If the applied voltage is large enough, then current flows between the electrodes, through the mixture, and an electrochemical reaction is induced.

According to the invention, at least two electrodes are provided to detect and measure the electric current. 15 When the reaction conditions are normal, the steady state current flowing between the electrodes lies within a predictable range of values. A steady state current which lies outside the expected range provides an indication that reaction condition within the apparatus are adverse. The 20 magnitude of the steady state current depends upon factors such as the electrode dimensions, the concentration of reagents in the reaction mixture, and on the voltage applied across the electrodes. The electrode current may also be affected by the presence of bubbles within the 25 apparatus.

A detection circuit is provided to compare the measured current with the expected current range, and to generate a signal indicating whether reaction conditions are normal, or adverse. In one embodiment, a single pair of 30 electrodes performs a dual function of both inducing the electrochemical reaction and allowing the current flowing through the mixture to be measured.

#### BRIEF DESCRIPTION OF THE FIGURES

Figure 1 is a schematic diagram of a microfluidic electrochemical reaction apparatus embodying the invention; Figure 2 is a schematic diagram of the detection circuit shown in Figure 1; and

5 Figure 3 is a graph showing the time variation of current measured by reaction chamber electrodes in the apparatus of Figure 1 for normal, and adverse reaction conditions.

#### DETAILED DESCRIPTION OF PREFERRED EMBODIMENTS

10 Figure 1 is a schematic diagram of a microfluidic electrochemical reaction apparatus embodying the invention. The apparatus 1 comprises a first mixing channel 3, second mixing channel 5, and reaction chamber 7. The first mixing channel 3, second mixing channel 5, and reaction chamber 7 are connected so that fluid may pass, in turn, through the first mixing channel 3, through the second mixing channel 5, and into the reaction chamber 7 via chamber inlet 8. The up-stream portion of the first mixing channel 3 is

15 connected to a substrate inlet 9, and an enzyme inlet 11 which supply the first mixing channel 3 respectively with substrate and enzyme. The up-stream portion of the second mixing channel 5 is connected to a mediator inlet 13 which supplies mediator to the second mixing channel 5. The reaction chamber 7 is connected to a waste outlet 15 which

20 allows fluid to pass out of the reaction chamber 7.

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The present invention relates particularly to electrochemical reaction chambers having dimensions in the micrometer scale. One exemplary use of a reaction chamber according to the present invention is as an analytical component on a miniature bio-chip.

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5           The substrate comprises molecules which may be electrochemically reacted with an enzyme in the reaction chamber 7. The substrate may comprise, for example, any one of a large array of xenobiotic compounds including drugs, pesticides and environmental pollutants. The enzyme may be any metabolising enzyme suitable for detoxifying the compound forming the substrate. Examples of such enzymes include proteins from the cytochrome P450 and flavin mono-oxygenase families. The mediator acts as a medium through 10 which electrons may flow, and may be provided by any suitable electrically conductive fluid. The purpose of the mediator is to allow electrons to be transferred from an electrode 17, located inside the reaction chamber 7, to the enzyme, so that an electrochemical reaction between the 15 enzyme and the substrate may be induced. The substrate, enzyme, and mediator are all provided in a fluid form so that they may flow along the mixing channels 3, 5.

20           In use, substrate is continually supplied to the first mixing channel 3 through substrate inlet 9 at a first predetermined flow rate, and enzyme is continually supplied to the first mixing channel 3 through enzyme inlet 11 at a second predetermined flow rate. The substrate and enzyme are combined in the first mixing channel 3, for example, by 25 diffuse mixing. The resulting substrate/enzyme mixture flows along the first mixing channel 3 to the second mixing channel 5. No reaction occurs between the substrate and enzyme in the first mixing channel 3 because no electrons are available to induce a reaction.

30           Mediator is continually supplied to the second mixing channel 5 through mediator inlet 13 at a third predetermined flow rate. Mediator is combined with the substrate/enzyme mixture in the second mixing channel 5, and the resulting substrate/enzyme/mediator mixture, hereinafter referred to as reagent mixture, flows through

the second mixing channel 5 to the reaction chamber 7 via chamber inlet 8. The proportions of substrate, enzyme, and mediator in the reagent mixture entering the reaction chamber 7 are determined by the relative magnitudes of the 5 first, second, and third flow rates, which may be adjusted to obtain a desired reagent mixture consistency. Reagent mixture is electrochemically reacted within the reaction chamber 7, and any unreacted reagent mixture, and reaction products are expelled through the waste outlet 15.

10 The reaction chamber 7 comprises two or more electrodes 17 which are located so that, in use, at least a portion of each electrode 17 is in electrical contact with the reagent mixture contained in the reaction chamber 7. The electrodes 17 are arranged to apply a predetermined 15 voltage across part of the region within the reaction chamber 7 through which reagent mixture passes, in use. The electrodes 17 are connected to a circuit arranged so that a first electrode, 17a for example, has a positive polarity, and a second electrode, 17b for example, has a negative polarity, thereby creating an electric field in 20 the region between the electrodes 17. Electric fields may also be created by several pairs of electrodes 17. When the 25 voltage applied across the electrodes 17 is larger than the redox potential of any of the reagents in the reagent mixture, electrons are caused to flow between the electrodes 17 in the form of an electric current through the mediator. Some of the electrons are transported from the electrode to enzyme molecules through the mediator, and an electrochemical reaction is induced between the enzyme 30 and the substrate.

The electric current flowing between the electrodes 17 may be measured by any suitable current measuring device connected to the electrodes 17. In the preferred embodiment, the electrodes 17 are coupled to a detection

circuit 19 arranged to detect and measure the current flow between the electrodes 17. Figure 2 is a schematic diagram of the detection circuit. The detection circuit 19 comprises inputs 21, a current detector 23, a comparator 25, a memory 27, and an output 31. The detection circuit 19 may be coupled to the electrodes 17 of the reaction chamber 7 via inputs 21 which receive current flowing between the electrodes 17. The inputs 21 are coupled to the current detector 23 which is arranged to detect and measure the current received by the inputs 21, and to generate a signal representative of the size of the current. The signal generated by the current detector 23 is received by the comparator 25 which is arranged to compare the measured current to a predetermined range of values defined by a signal received by the comparator 25 from the memory 27. The memory 27 is arranged to store one or more predetermined ranges, for example, for a steady state current or for a maximum current. The comparator 25 is further arranged to generate a signal indicating whether the measured current lies inside, or outside a selected range of values stored in the memory 27. The signal generated by the comparator 25 is output from the detection circuit 19 via output 31. When the detection circuit 19 is coupled to the electrodes 17, The output 31 provides an indication of whether the electrode current lies inside or outside a particular predetermined range.

In the preferred embodiment, the electrodes 17 provided to induce the electrochemical reaction, and the electrodes 17 provided to measure the current flow through the reagent mixture are the same electrodes 17. In this embodiment, the electrodes 17 perform a dual function, and the components required to work the invention are advantageously provided by existing components.

Alternatively, separate pairs of electrodes 17 may be provided for each individual function.

Figure 3 is a graph showing the time variation of the current flowing between the electrodes 17 during use of the apparatus 1. Initially, the first mixing channel 3, second mixing channel 5, and reaction chamber 7 are filled with air. Since air is electrically insulating, if the voltage applied across the electrodes 17 is below a particular threshold, then no current will flow between the electrodes 17. From time zero, mediator is continually introduced into the second mixing channel 5, which then flows along the second mixing channel 5 towards the reaction chamber 7, displacing air from the second mixing channel 5. During a first time period 43, mediator has not yet reached the reaction chamber 7, and during this time, the electrode current is zero because the reaction chamber is filled with air.

Following the first time period 43, mediator flowing along the second mixing channel 5 arrives at the reaction chamber 7, and subsequently fills the reaction chamber 7. Since the mediator is electrically conductive, the presence of mediator within the reaction chamber 7 allows current to flow between the electrodes 17. During the time the mediator is gradually filling the reaction chamber 7, the electrode current rises from a zero value to a predictable maximum value, which is dependant on the electrical conductivity of the mediator and other factors. During a second time period 45, immediately following the first time period 43, the reaction chamber 7 is filled with mediator, and the electrode current maintains the maximum value.

From time zero, substrate and enzyme are continually introduced into the first mixing channel 3, which then flows along the first mixing channel 3 towards the reaction chamber 7, displacing air from the first mixing channel 3.

During the first time period 43 at least, the air displaced from the first mixing channel 3 forms one or more air pockets between the substrate/enzyme mixture flowing along the first mixing channel 3, and the mediator flowing along the second mixing channel 5. As the substrate/enzyme mixture, and mediator flow along the mixing channels 3, 5, the air pockets are transported within the fluid flow to the reaction chamber 7. When the air pockets arrive at the reaction chamber 7, the mediator which previously filled the reaction chamber 7 during the second time period 45 is displaced, and gradually replaced, by air forming the air pockets. During the time air is gradually filling the reaction chamber 7, the electrode current decreases from the maximum value when the reaction chamber is filled with mediator, to a zero value when the reaction chamber is filled with air. During a third time period 47, immediately following the second time period 45, the reaction chamber 7 is filled with air, and the electrode current is maintained at a zero value.

Since substrate, enzyme, and mediator are being continuously introduced into the mixing channels 3, 5, after a particular time, the substrate/enzyme mixture flowing along the first mixing channel 3 will eventually meet and combine with the mediator flowing along the second mixing channel 5 to form a mediator/enzyme/substrate mixture in the second mixing channel 5. It can be seen that this reagent mixture first begins to form upstream from, and adjacent to the air pockets. When the reagent mixture upstream from, and adjacent to the air pockets arrives at, and begins to fill the reaction chamber 7, the electrode current increases from a zero value, up to a particular value. This value depends on the electrical conductivity of the reagent mixture, which in turn, is dependant on the consistency of the mixture. During a forth time period 49,

immediately following the third time period 47, the reaction chamber 7 is filled with reagent mixture, and the electrode current maintains a steady state value. At subsequent times, reagent mixture flows continuously through the reaction chamber 7, and the electrode current is maintained at an approximately constant steady state value.

It can be seen that the variation in the electrode current during the first 43, second 45, and third 47 time periods is caused by the presence of air within the apparatus 1, and also the presence of substrate, enzyme, and mediator which are not fully mixed. This variation may be viewed as transient variation in the electrode current. During the forth time period 49 however, the conditions within the apparatus 1 have reached a steady state, and it is during this time period that the electrochemical reaction is initiated. It is important that sufficient time is allowed for the transient conditions to pass, and for the steady state condition to be maintained. Once the acceptable steady state electrode current has been maintained for at least a predetermined time, the voltage applied across the electrodes 17 may be controlled in order to induce an electrochemical reaction. The precise timing of the sequence described above may be controlled by altering the geometry of the apparatus 1.

The steady state electrode current of the forth time period 49 is lower than the maximum electrode current of the second time period 45 because the presence of substrate and enzyme in the reaction chamber 7 lowers the electrical conductivity of the reagent mixture. When an electrochemical reaction is induced, the consumption of electrons by the reaction results in a further reduction in electrode current. In the preferred embodiment, the steady state electrode current is approximately 50% of the

maximum electrode current. The steady state current may fluctuate around an average value, but under normal reaction conditions, remains within a predetermined tolerance range.

5 The steady state electrode current depends on factors including the dimensions of the electrodes 17, the voltage applied across the electrodes 17, and the concentrations of substrate, enzyme, and mediator in the reagent mixture. If the value of each of these factors is determined, then 10 the steady state electrode current is also determined and predictable. It is the predictability of the steady state current, that allows adverse reaction conditions to be detected when the steady state current does not correspond to the predicted current.

15 It is important that during an electrochemical reaction, the conditions within the reaction chamber 7 are correct, otherwise the reaction process may be affected, and any analysis carried out relating to the reaction may give erroneous results. Some factors that may affect a 20 reaction include the concentrations of substrate, enzyme, and mediator in the reagent mixture, and the surface dimensions of the electrodes 17. Other factors include the electrode 17 spacing, and the voltage applied across the electrodes 17. Another way that reaction conditions may be 25 affected is by the presence of one or more bubble within the apparatus 1.

Since the apparatus 1 is initially filled with air, there is a risk that an air bubble could be trapped within the fluid flowing through the apparatus 1. In one case, a 30 bubble may become trapped in one of the inlets 9, 11, 13, restricting the flow rate of substrate, enzyme, or mediator, and thereby altering the constitution of the reaction mixture. In another case, a bubble may become trapped in one of the mixing channels 3, 5, thereby

affecting the mixing of the constituents resulting in non-uniform mixture consistency. A bubble trapped in a mixing channel 3, 5, or any other blockage, may also affect the flow within the apparatus 1, or even prevent flow completely. In yet a further case, a bubble may become trapped inside the reaction chamber 7, and locate on the surface of one or more of the electrodes 17. In this case, the effective surface area of the electrode 17 is reduced, resulting in an increase in electrical resistance across the electrodes 17, and a reduction in electrode current. The reaction conditions may also be affected even if a bubble trapped inside the reaction chamber 7 is not located on the surface of an electrode 17.

The graph shown as a dotted line in Figure 3 shows the time variation of the electrode current in the case when a bubble is trapped inside the reaction chamber 7, and located on a surface of one of the electrodes 17. During the second 45 and forth 49 time periods, the presence of a bubble on the electrode 17 increases the electrical resistance between the electrodes 17, and causes a reduction in electrode current relative to the case when no bubble is present. The presence of the bubble causes the electrode current to lie outside the range in which the current would be expected to lie if the conditions in the apparatus 1 were normal.

According to the invention, a deviation of the steady state electrode current from the expected value provides an indication that the reaction conditions within the apparatus 1 are adverse. Conversely, an electrode current which conforms to an expected value provides an indication that reaction conditions are normal. Accordingly, any analysis of a reaction may be accepted or rejected according to the output signal generated by the detection circuit 19, which indicates whether the electrode current

lies inside, or outside an expected range of values. In one embodiment, the detection circuit 19 generates a signal indicating normal reaction conditions only if the steady state electrode current lies within the predetermined range of values for a predetermined time. If the electrode current falls outside the predetermined range during this time, then a signal is generated to indicate adverse reaction conditions. In this way, it can be seen that the electrodes 17 act as a sensor to enable detection of adverse reaction conditions.

In one embodiment, the value of the maximum electrode current during the second time period 45 is also monitored. In this case, during the second time period 45, the electrode current is measured by the detection circuit 19, and compared to the expected maximum electrode current. If the electrode current lies outside a predetermined range of values, then the detection circuit 19 generates a signal to indicate adverse condition present within the apparatus 1. In this way, if the detection circuit 19 determines during the second time period 45 that conditions within the apparatus 1 are adverse, then the process may be terminated early, saving time, and volume of substrate, enzyme, and mediator used.

It is understood that a deviation of electrode current from the expected value is indicative of adverse reaction conditions which may be caused by any number of factors, of which the presence of a bubble within the apparatus is merely one example. A change in concentrations of substrate, enzyme, or mediator, for example, may also affect the reaction conditions. If the concentration of mediator is higher than normal, then the electrical conductivity of the reagent mixture is higher than normal. Accordingly, the electrode current is higher than would be expected if the mediator concentration were normal. If the

mediator concentration deviates sufficiently from the desired concentration, then the steady state electrode current will lie outside the expected range. In a further example, an incorrect voltage applied across the electrodes 5 17 will cause a deviation in electrode current which may then be detected.

It will be appreciated by the skilled person that the invention is applicable to both static and dynamic electrochemical reaction chambers. In the case of static 10 electrochemical reaction chambers, fluid flow may be discontinuous, for example if a batch process is employed, or there may be no flow of fluid at all. In this case, the presence of bubbles in the reaction chamber 7 causes a characteristic decrease in electrode current over time, 15 indicating adverse reaction conditions. In the case of a dynamic electrochemical chamber, where fluid flow through the reaction chamber 7 is continuous, bubbles passing through the reaction chamber 7 cause brief transient reductions in the electrode current, which may be 20 continually monitored.

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CLAIMS

1. A microelectrochemical reaction chamber comprising an inlet for allowing passage of electrochemical reaction constituents into the reaction chamber, and at least two electrodes, the electrodes arranged to have, in use, electrical contact with the reaction fluid, the electrodes being connectable to a detection circuit arranged to detect a current flow between the electrodes, and arranged to generate a signal if the current falls outside a predetermined range, thereby indicating adverse reaction conditions.
2. A microelectrochemical reaction chamber according to claim 1 wherein at least two of the electrodes are arranged to apply a voltage across at least a part of the region through which the reaction constituents pass, in use, thereby inducing an electrochemical reaction between the constituents.
3. A microelectrochemical reaction chamber according to claim 2 wherein the electrodes used to induce the electrochemical reaction are the same as the electrodes used to detect a current flow.
4. A microelectrochemical reaction chamber according to claim 1 wherein the chemical reaction constituents form a mixture including a substrate, an enzyme, and a mediator.
5. A microelectrochemical reaction chamber according to claim 4 wherein the substrate comprises a xenobiotic compound.

6. A microelectrochemical reaction chamber according to claim 4 wherein the enzyme comprises a protein from the cytochrome P450 or flavin mono-oxygenase families.
- 5 7. A method of determining whether the reactions conditions in a microelectrochemical reaction chamber are correct comprising feeding reaction fluids in turn into a reaction chamber, monitoring the current flow between electrodes in the reaction chamber, determining whether the current flow lies outside a first defined range, and, if the current lies outside the first defined range, determining that the reaction conditions in the chamber are adverse.
- 10 8. The method according to claim 7 wherein the step of feeding reaction fluids into the chamber further comprises the steps of feeding mediator into the chamber, and subsequently feeding a substrate and an enzyme into the chamber.

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- 15 9. The method according to claim 8 wherein the step of feeding mediator into the chamber further comprises the steps of monitoring the current flow between the electrodes when the reaction chamber contains mediator only, determining whether the current flow lies outside a second defined range, and, if the current lies outside the second defined range, determining that the reaction conditions in the chamber are adverse.
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10. The method according to claim 7 wherein the step of feeding reaction fluids into the chamber further comprises the step of feeding a mixture of substrate, enzyme, and mediator into the reaction chamber.
- 5 11. The method according to claim 7 wherein the step of determining that the reaction conditions in the chamber are adverse further includes the step of determining that a bubble is located within the reaction chamber.
- 10 12. The method according to claim 7 wherein the step of determining that the reaction conditions in the chamber are adverse further includes the step of determining that the concentrations of reaction fluids within the reaction chamber are incorrect.
- 15 13. The method according to claim 7 wherein the step of determining that the reaction conditions in the chamber are adverse further includes the step of determining that a voltage applied across the electrodes is incorrect.
- 20 14. The method according to claim 7 further comprising the step of generating a signal to indicate whether the reaction conditions in the chamber are adverse.
- 25 15. The method according to claim 7 wherein the step of feeding reaction fluids into the chamber further comprises the step of feeding each reaction fluid into the chamber at a predetermined flow rate.

16. Apparatus for determining that the reaction conditions for a microelectrochemical reaction are adverse comprising a microelectrochemical reaction chamber and a detection circuit, wherein the microelectrochemical reaction chamber comprises:

- 5 - an inlet for allowing passage of electrochemical reaction constituents into the reaction chamber; and
- 10 - at least two electrodes, the electrodes arranged to have, in use, electrical contact with the reaction fluid, the electrodes being connected to the detection circuit;

15 and wherein the detection circuit comprises:

- 20 - a current detector arranged to detect and measure a current flow between the electrodes of the reaction chamber;
- a comparator for comparing the current to a predetermined range, and arranged to generate a signal for indicating whether the current lies outside the predetermined range, thereby indicating adverse reaction conditions.

## ABSTRACT

A method and apparatus for detecting that reaction conditions in a microelectrochemical reaction chamber are adverse are disclosed. The reaction chamber comprises 5 electrodes arranged to pass an electric current through reaction mixture located within the reaction chamber, thereby inducing an electrochemical reaction. A detection circuit is provided to detect and measure the electric current flowing between the electrodes. The detection 10 circuit is arranged to generate a signal indicating whether the measured current lies inside or outside a predetermined range of values. If the measured current lies outside the expected range of values, then the reaction conditions are adverse. A single pair of electrodes may perform a dual 15 function of both inducing the electrochemical reaction and allowing the current flowing through the mixture to be measured.

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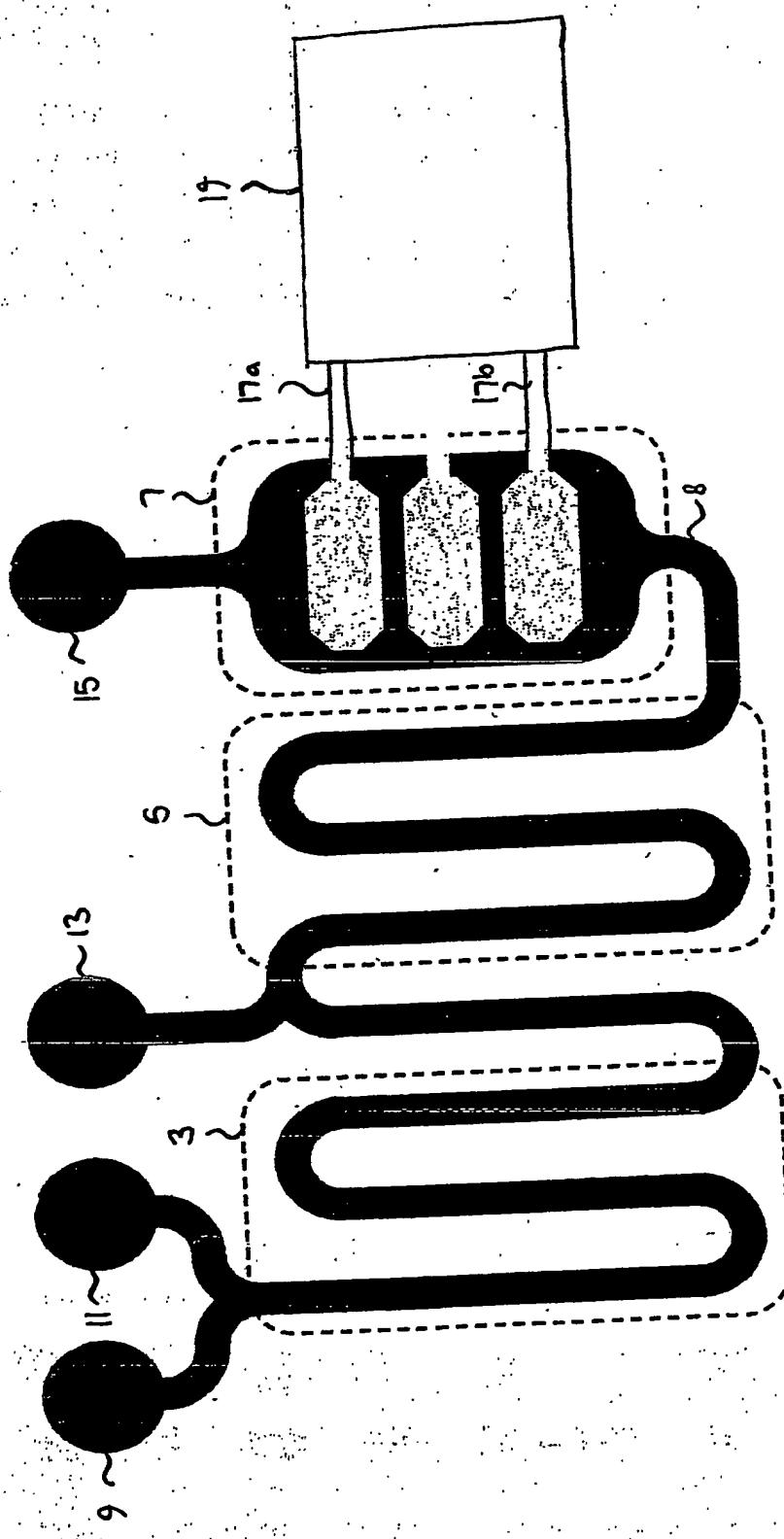
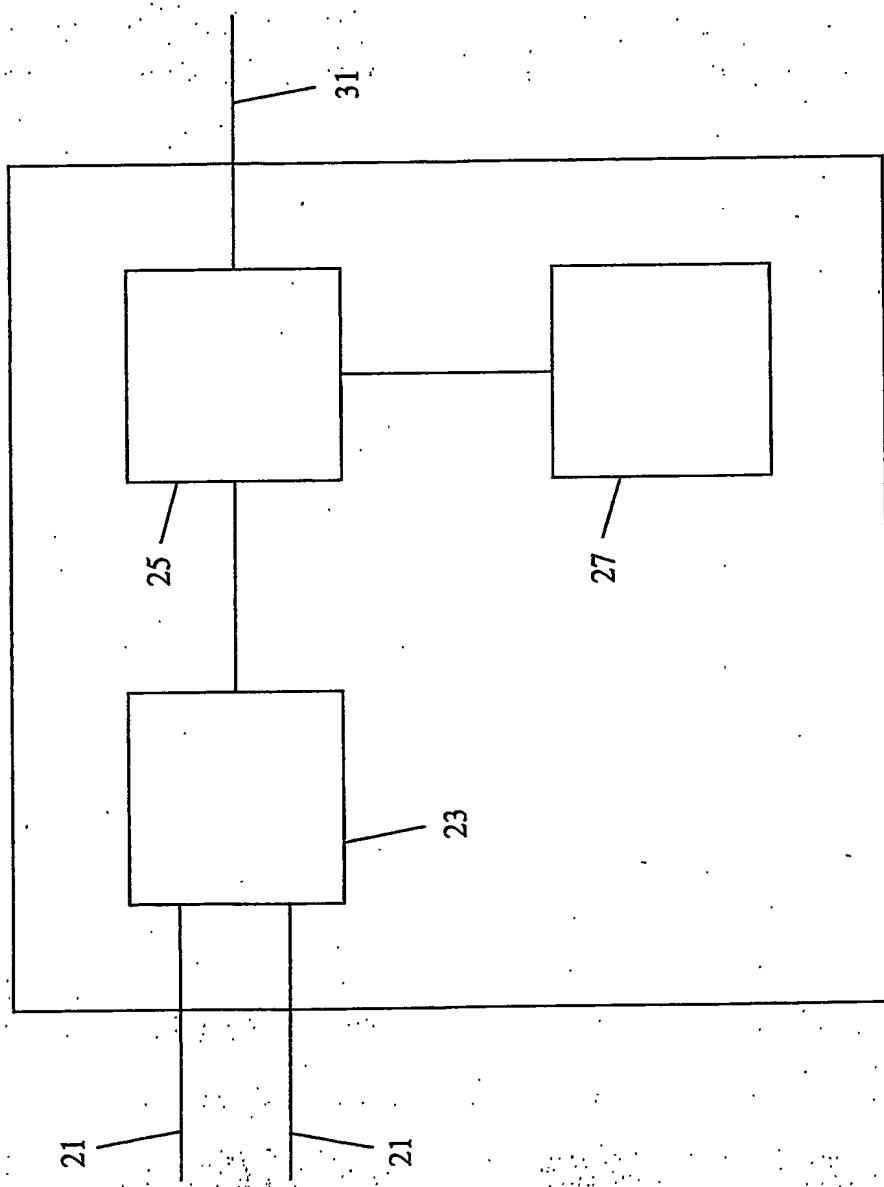


Figure 1

19

Figure 2



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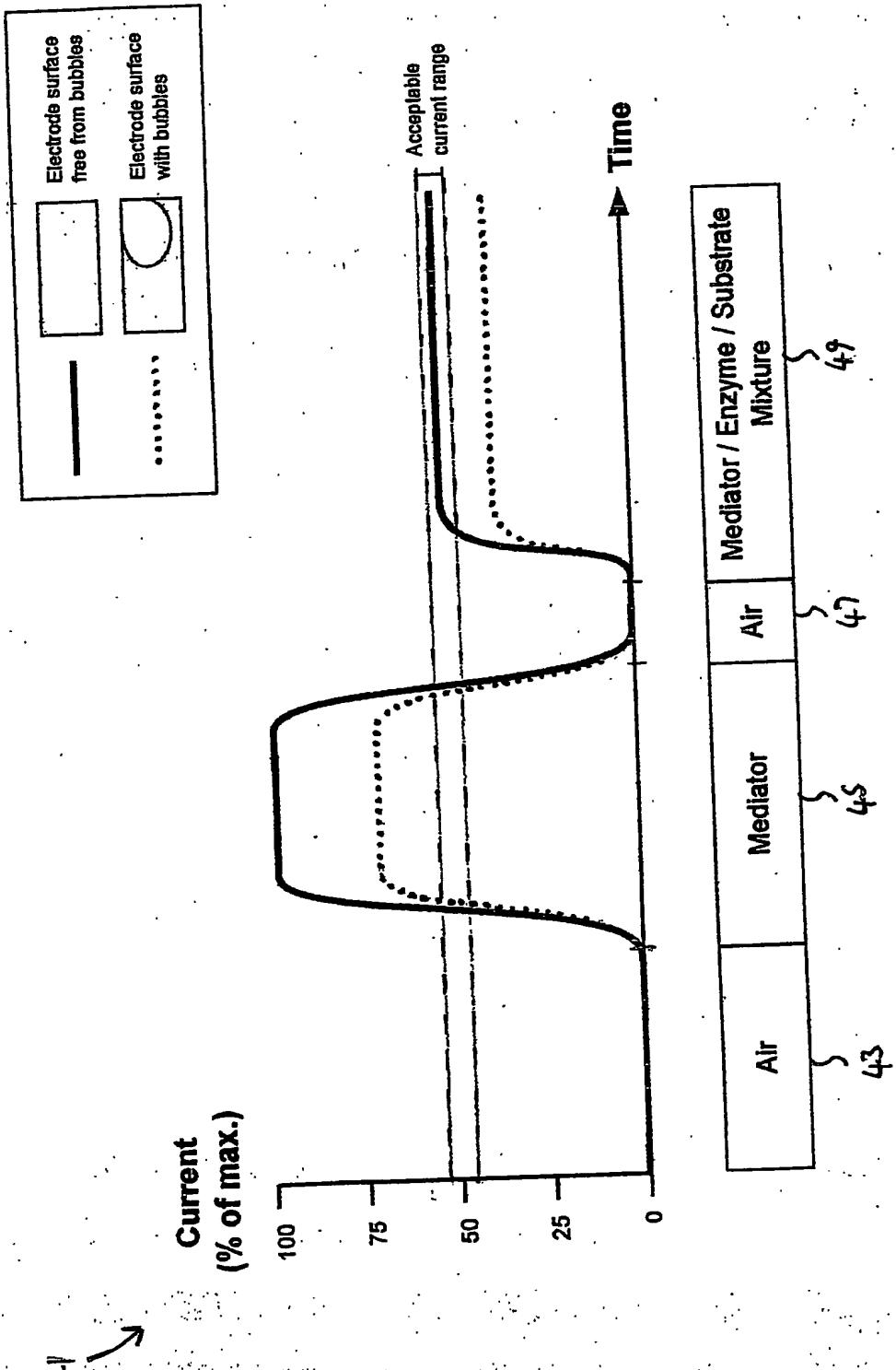


Figure 3